

Instructions: You should have with you several number two pencils, an eraser, your 3" x 5" note card, a calculator, and your University ID Card. If you have notes or a phone with you, place them in a sealed backpack and place the backpack OUT OF SIGHT or place the notes directly on the table at the front of the room.

Fill in the front page of the Scantron sheet with your last name, first name, middle initial, student identification number, and section number (below). Leave the test form number blank.

- Section 001 (MWF 8am with Dr. Nafshun)
- Section 002 (MWF 9am with Dr. Nafshun)
- Section 003 (MWF 10am with Dr. Nafshun)
- Section 004 (MWF 11am with Dr. Watson)
- Section 005 (MWF 1pm with Dr. Nyman)
- Section 006 (MWF 2pm with Dr. Barth)
- Section 007 (MWF 3pm with Dr. Burand)

This exam consists of 32 multiple-choice questions; each has 5 points attached. When you finish this exam, proceed to the proctor. Flash your OSU ID Card and submit your completed Scantron form. You may take your notecard and this exam packet with you.

1 1A																	18 8A
1 H Hydrogen 1.008												13 3A	14 4A	15 5A	16 6A	17 7A	2 He Helium 4.003
3 Li Lithium 6.941	4 Be Beryllium 9.012											5 B Boron 10.81	6 C Carbon 12.01	7 N Nitrogen 14.01	8 O Oxygen 16.00	9 F Fluorine 19.00	10 Ne Neon 20.18
11 Na Sodium 22.99	12 Mg Magnesium 24.31	3 3B	4 4B	5 5B	6 6B	7 7B	8 8B	9 8B	10	11 1B	12 2B	13 Al Aluminum 26.98	14 Si Silicon 28.09	15 P Phosphorus 30.97	16 S Sulfur 32.07	17 Cl Chlorine 35.45	18 Ar Argon 39.95
19 K Potassium 39.10	20 Ca Calcium 40.08	21 Sc Scandium 44.96	22 Ti Titanium 47.87	23 V Vanadium 50.94	24 Cr Chromium 52.00	25 Mn Manganese 54.94	26 Fe Iron 55.85	27 Co Cobalt 58.93	28 Ni Nickel 58.69	29 Cu Copper 63.55	30 Zn Zinc 65.38	31 Ga Gallium 69.72	32 Ge Germanium 72.64	33 As Arsenic 74.92	34 Se Selenium 78.96	35 Br Bromine 79.90	36 Kr Krypton 83.80
37 Rb Rubidium 85.47	38 Sr Strontium 87.62	39 Y Yttrium 88.91	40 Zr Zirconium 91.22	41 Nb Niobium 92.91	42 Mo Molybdenum 95.96	43 Tc Technetium (98)	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.91	46 Pd Palladium 106.42	47 Ag Silver 107.87	48 Cd Cadmium 112.41	49 In Indium 114.82	50 Sn Tin 118.7	51 Sb Antimony 121.76	52 Te Tellurium 127.60	53 I Iodine 126.90	54 Xe Xenon 131.29
55 Cs Cesium 132.91	56 Ba Barium 137.33	57 La Lanthanum 138.91	72 Hf Hafnium 178.49	73 Ta Tantalum 180.95	74 W Tungsten 183.84	75 Re Rhenium 186.21	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.97	80 Hg Mercury 200.59	81 Tl Thallium 204.38	82 Pb Lead 207.2	83 Bi Bismuth 208.98	84 Po Polonium (208.98)	85 At Astatine (209.99)	86 Rn Radon (222.02)
87 Fr Francium (223.02)	88 Ra Radium (226.03)	89 Ac Actinium (227.03)	104 Rf Rutherfordium (261.11)	105 Db Dubnium (262.11)	106 Sg Seaborgium (266.12)	107 Bh Bohrium (264.12)	108 Hs Hassium (269.13)	109 Mt Meitnerium (268.14)	110 Ds Darmstadtium (271)	111 Rg Roentgenium (272)	112 Cn Copernicium (285)	113	114 Fl Flerovium (289)	115	116 Lv Livermorium (293)	117	118

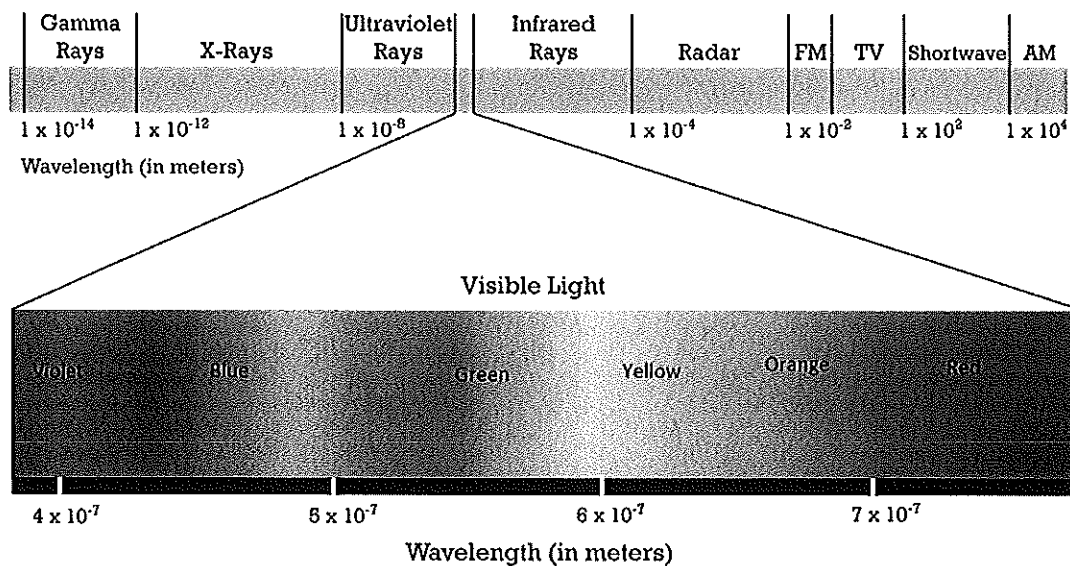
58 Ce Cerium 140.12	59 Pr Praseodymium 140.91	60 Nd Neodymium 144.24	61 Pm Promethium (145)	62 Sm Samarium 150.36	63 Eu Europium 151.96	64 Gd Gadolinium 157.25	65 Tb Terbium 158.93	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93	68 Er Erbium 167.26	69 Tm Thulium 168.93	70 Yb Ytterbium 173.05	71 Lu Lutetium 174.97
90 Th Thorium 232.04	91 Pa Protactinium 231.04	92 U Uranium 238.03	93 Np Neptunium (237.05)	94 Pu Plutonium (244.06)	95 Am Americium (243.06)	96 Cm Curium (247.07)	97 Bk Berkelium (247.07)	98 Cf Californium (251.08)	99 Es Einsteinium (252.08)	100 Fm Fermium (257.10)	101 Md Mendelevium (258.10)	102 No Nobelium (259.10)	103 Lr Lawrencium (262.11)

TABLE 1.2 SI Prefix Multipliers

Prefix	Symbol	Multiplier	
exa	E	1,000,000,000,000,000,000	(10 ¹⁸)
peta	P	1,000,000,000,000,000	(10 ¹⁵)
tera	T	1,000,000,000,000	(10 ¹²)
giga	G	1,000,000,000	(10 ⁹)
mega	M	1,000,000	(10 ⁶)
kilo	k	1000	(10 ³)
deci	d	0.1	(10 ⁻¹)
centi	c	0.01	(10 ⁻²)
milli	m	0.001	(10 ⁻³)
micro	μ	0.000001	(10 ⁻⁶)
nano	n	0.000000001	(10 ⁻⁹)
pico	p	0.000000000001	(10 ⁻¹²)
femto	f	0.000000000000001	(10 ⁻¹⁵)
atto	a	0.000000000000000001	(10 ⁻¹⁸)

1 inch = 2.54 cm (exact)	1 mile = 1.60934 km
1000 mL = 1 L	1 mL = 1 cm ³
1 mole (N _A) = 6.022 x 10 ²³	c = 3.00 x 10 ⁸ m/s
h = 6.626 x 10 ⁻³⁴ J·s	1 m = 1 x 10 ⁹ nm

9 Polyatomic Ions		Material	Specific Heat (J/g°C)
Name	Formula		
Hydroxide	OH ⁻	H ₂ O (l)	4.184
Cyanide	CN ⁻	CH ₃ CH ₂ OH (l)	2.42
Nitrate	NO ₃ ⁻	Pb (s)	0.128
Acetate	CH ₃ COO ⁻	Au (s)	0.128
Carbonate	CO ₃ ²⁻	Ag (s)	0.235
Phosphate	PO ₄ ³⁻	Cu (s)	0.385
Hydronium	H ₃ O ⁺	Fe (s)	0.449
Ammonium	NH ₄ ⁺	Al (s)	0.903
Sulfate	SO ₄ ²⁻	Sand	0.84
		Granite	0.790



$\lambda = \frac{h}{mv}$	$q = mc\Delta T$	$q = m\Delta H$
$\Delta E = q + w$	$q = c \Delta T$	$c = 3.00 \times 10^8 \text{ m/s}$
$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$	$\nu = \frac{c}{\lambda}$	$E = h\nu$
$\frac{1}{\lambda} = R_H \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$	$R_H = 2.180 \times 10^{-18} \text{ J/photon}$	$R_H = 10973731.6 \text{ m}^{-1}$
$1 \text{ J} = 1 \frac{\text{kg} \cdot \text{m}^2}{\text{s}^2}$	electron mass = $9.10938 \times 10^{-31} \text{ kg}$	$E = mc^2$
$E\Psi = H\Psi$	$\Delta E = q + w$	$W = -P\Delta V$
$\Delta H = \Delta E - P\Delta V$	$P(r) = 4\pi r^2 \Psi^2$	$(\Delta x)(m\Delta v) \geq h/4\pi$
$E_n = -hcR_\infty/n^2$	$E_n = -RZ^2/n^2$	

1. A student combusts a compound in a bomb calorimeter, and the temperature of the calorimeter increases from 25.62 °C to 27.42 °C. The calorimeter has a constant of 91.82 J/°C. How many significant figures can the student report for the ΔE in this reaction?

a) 1

b) 2

c) 3

d) 4

e) 5

$$\begin{aligned} q &= \left(91.82 \frac{\text{J}}{^\circ\text{C}} \right) (27.42^\circ\text{C} - 25.62^\circ\text{C}) \\ &= \underbrace{(91.82 \text{ J/}^\circ\text{C})}_4 \underbrace{(1.80^\circ\text{C})}_3 \\ &= \frac{165.}{3} \end{aligned}$$

$$\frac{165.}{3}$$

2. A solid cube is measured to be 0.451 inches on each side. The mass of the cube is 12.9 g. Which of the following metals is this cube most likely to be?

a) Yttrium (4.45 g/cm³)

b) Lead (11.3 g/cm³)

c) Iridium (22.4 g/cm³)

d) Tellurium (4.95 g/cm³)

e) Niobium (8.58 g/cm³)

$$s = 0.451 \text{ in}$$

$$s = 0.451 \text{ in} \left(\frac{2.54 \text{ cm}}{1 \text{ in}} \right) = 1.15 \text{ cm}$$

$$V = s^3 = (1.15 \text{ cm})^3 = 1.50 \text{ cm}^3$$

$$d = \frac{m}{V} = \frac{12.9 \text{ g}}{1.50 \text{ cm}^3} = 8.58 \text{ g/cm}^3$$

3. Which of the following statements about subatomic particles is **TRUE**?

a) A neutral atom contains the same number of protons and electrons.

b) Protons have about the same mass as electrons. ~~X~~

c) Electrons make up most of the mass of an atom. ~~X~~

d) Protons and neutrons have opposite, but equal in magnitude, charges. ~~X~~

e) Neutrons and electrons are found in the nucleus of an atom. ~~X~~

4. How many moles of Kr are contained in 398 mg of Kr?

- a) 4.75×10^{-3} moles Kr
- b) 33.4 moles Kr
- c) 2.11×10^{-4} moles Kr
- d) 2.99×10^{-3} moles Kr
- e) 1.19×10^{-4} moles Kr

$$398 \times 10^{-3} \text{ g Kr} \left(\frac{1 \text{ mol}}{83.8 \text{ g}} \right) = 4.75 \times 10^{-3} \text{ mol Kr}$$

5. The fictional element Beaverium (Bv) has two stable isotopes: ^{145}Bv (mass = 144.9362 amu) and ^{148}Bv (mass = 147.9177 amu). Which of the following could not be the average atomic mass of Beaverium?

- a) 145.7731 amu
- b) 144.9968 amu
- c) 146.3214 amu
- d) 147.9882 amu

can not be greater than 147.9177 amu - must be between 144.9362 & 147.9177 amu

e) any of these could be the average atomic mass based on the information given

6. How many total electrons are present in 0.1415 g of oxide ions?

- a) 2.23×10^{23} electrons
- b) 3.20×10^{22} electrons
- c) 4.26×10^{22} electrons
- d) 8.84×10^{22} electrons
- e) 5.33×10^{22} electrons

$\text{O}^{2-} \rightarrow \text{O}^{2-} \text{ } 10e^- \text{ each}$

$$0.1415 \text{ g O}^{2-} \left(\frac{1 \text{ mol}}{16.00 \text{ g}} \right)$$
$$= 0.00884 \text{ mol O}^{2-}$$
$$0.00884 \text{ mol O}^{2-} \left(\frac{6.022 \times 10^{23} \text{ O}^{2-}}{1 \text{ mol}} \right) =$$
$$= 5.33 \times 10^{21} \text{ O}^{2-} \text{ ions}$$
$$5.33 \times 10^{21} \text{ O}^{2-} \text{ ions} \left(\frac{10 e^-}{1 \text{ O}^{2-}} \right) = 5.33 \times 10^{22} e^-$$

7. Which of the following pairs of formulas and names are incorrect?

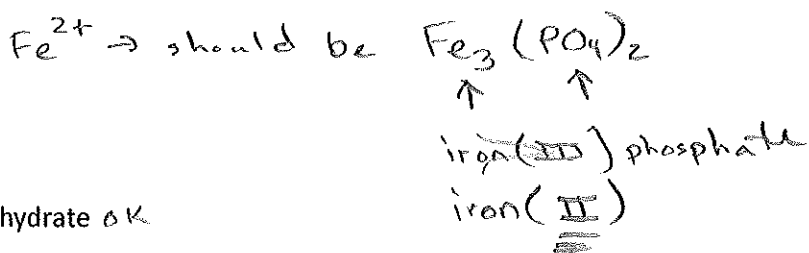
a) $\text{Fe}_3(\text{PO}_4)_2$; iron (III) phosphate

b) NH_4OH ; ammonium hydroxide OK

c) CaCl_2 ; calcium chloride OK

d) $\text{Al}(\text{NO}_3)_3 \cdot 6\text{H}_2\text{O}$; aluminum nitrate hexahydrate OK

e) TiS_2 ; titanium (IV) sulfide OK



8. How many grams of water are in 16.00 g of $\text{FeSO}_4 \cdot 4\text{H}_2\text{O}$? Molar Mass of $\text{FeSO}_4 \cdot 4\text{H}_2\text{O} = 224.0 \text{ g/mol}$.

a) 5.149

b) 4.571

c) 6.429

d) 3.502

e) 4.857

$$16.00 \text{ g } \text{FeSO}_4 \cdot 4\text{H}_2\text{O} \left(\frac{1 \text{ mol}}{224.0 \text{ g}} \right) = 0.07143 \text{ mol } \text{FeSO}_4 \cdot 4\text{H}_2\text{O}$$

$$0.07143 \text{ mol } \text{FeSO}_4 \cdot 4\text{H}_2\text{O} \left(\frac{4 \text{ mol } \text{H}_2\text{O}}{1 \text{ mol } \text{FeSO}_4 \cdot 4\text{H}_2\text{O}} \right) = 0.2857 \text{ mol } \text{H}_2\text{O}$$

$$0.2857 \text{ mol } \text{H}_2\text{O} \left(\frac{18.02 \text{ g}}{1 \text{ mol}} \right) = 5.149 \text{ g } \text{H}_2\text{O}$$

9. Aqueous iron(II) chloride reacts with aqueous sodium phosphate to produce solid iron(II) phosphate and aqueous sodium chloride. Which represents a balanced equation of this?

a) $3 \text{FeCl}_2(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow \text{Fe}_2(\text{PO}_4)_2(\text{s}) + 3 \text{NaCl}(\text{aq})$

b) $\text{FeCl}_2(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow \text{Fe}_3(\text{PO}_4)_2(\text{s}) + \text{NaCl}(\text{aq})$

c) $\text{FeCl}_3(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow \text{Fe}_2\text{PO}_4(\text{s}) + 3 \text{NaCl}(\text{aq})$

d) $3 \text{FeCl}_2(\text{aq}) + 2 \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow \text{Fe}_3(\text{PO}_4)_2(\text{s}) + 6 \text{NaCl}(\text{aq})$

e) $3 \text{FeCl}_2(\text{aq}) + \text{Na}_3\text{PO}_4(\text{aq}) \rightarrow 2 \text{Fe}_3(\text{PO}_4)_2(\text{s}) + 3 \text{NaCl}(\text{aq})$



10. 8.84 g of a compound is analyzed and found to contain 3.53 g of C, 0.295 g of H, 4.08 g of N, and 0.935 g of O. Which of the following could be the molecular formula for this compound?

a) $C_{10}H_{10}N_{10}O_2$ *molecular*

b) $C_4H_8N_6O_2$

c) $C_2H_4N_6O$

d) $C_5H_{10}N_3O_3$

e) $C_8H_8N_8O_4$

$$3.53 \text{ g C} \left(\frac{1 \text{ mol}}{12.01 \text{ g}} \right) = 0.294 \text{ mol C} / 0.0584$$

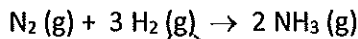
$$0.295 \text{ g H} \left(\frac{1 \text{ mol}}{1.008 \text{ g}} \right) = 0.293 \text{ mol H} / 0.0584$$

$$4.08 \text{ g N} \left(\frac{1 \text{ mol}}{14.01 \text{ g}} \right) = 0.291 \text{ mol N} / 0.0584$$

$$0.935 \text{ g O} \left(\frac{1 \text{ mol}}{16.00 \text{ g}} \right) = 0.0584 \text{ mol O} / 0.0584$$

$\times 2$ $C_5H_5N_5O$
empirical

11. Consider the following reaction:



What mass of hydrogen gas must be reacted in excess nitrogen gas to give a theoretical yield of 75.0 g of NH_3 gas?

a) 6.60 g H_2

b) 13.3 g H_2

c) 46.5 g H_2

d) 11.3 g H_2

e) 5.92 g H_2

① $75.0 \text{ g } NH_3 \left(\frac{1 \text{ mol}}{17.03 \text{ g}} \right) = 4.40 \text{ mol } NH_3$

② $4.40 \text{ mol } NH_3 \left(\frac{3 \text{ mol } H_2}{2 \text{ mol } NH_3} \right) = 6.60 \text{ mol } H_2$

③ $6.60 \text{ mol } H_2 \left(\frac{2.02 \text{ g}}{1 \text{ mol}} \right) = 13.3 \text{ g } H_2$

12. Indium can react with oxygen gas to form indium(III) oxide. If this reaction has a percent yield of 63.8%, how many mol of indium must be reacted in excess oxygen to yield 3.00 mol of indium(III) oxide?

a) 5.74 mol In

b) 4.70 mol In

c) 9.40 mol In

d) 3.00 mol In

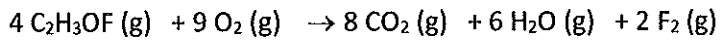
e) 6.00 mol In

$$2 In(s) + \frac{3}{2} O_2(g) \rightarrow In_2O_3$$

6.00 mol ← 3.00 mol

$$\frac{6.00 \text{ mol}}{0.638} = 9.40 \text{ mol In}$$

13. Consider the following reaction:



When 5.82 mol of $\text{C}_2\text{H}_3\text{OF}$ and 10.35 mol of O_2 are initially present in the reaction mixture, how much of which reactant will remain after the reaction goes to completion?

a) 1.15 mol O_2 remain

b) 4.53 mol $\text{C}_2\text{H}_3\text{OF}$ remain

c) 1.22 mol $\text{C}_2\text{H}_3\text{OF}$ remain

d) 2.97 mol $\text{C}_2\text{H}_3\text{OF}$ remain

e) 2.75 mol O_2 remain

$$5.82 \text{ mol } \text{C}_2\text{H}_3\text{OF} \left(\frac{9 \text{ mol } \text{O}_2}{4 \text{ mol } \text{C}_2\text{H}_3\text{OF}} \right) = 13.09 \text{ mol } \text{O}_2 \text{ would be used} - \text{O}_2 \text{ is L.R.}$$

$$10.35 \text{ mol } \text{O}_2 \left(\frac{4 \text{ mol } \text{C}_2\text{H}_3\text{OF}}{9 \text{ mol } \text{O}_2} \right) = 4.60 \text{ mol } \text{C}_2\text{H}_3\text{OF} \text{ will be used}$$

$$5.82 \text{ mol} - 4.60 \text{ mol} = 1.22 \text{ mol } \text{C}_2\text{H}_3\text{OF} \text{ remain}$$

14. What mass of stones ($C_s = 0.841 \text{ J/g}\cdot^\circ\text{C}$) that are heated to 815.0°C must be added to 750.0 g of water ($C_s = 4.184 \text{ J/g}\cdot^\circ\text{C}$) initially at 22.3°C to raise the temperature of the water to the boiling point (100.0°C)?

a) 793 g stones

b) 405 g stones

c) 445 g stones

d) 366 g stones

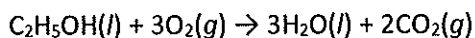
e) 557 g stones

$$q_{\text{stones}} = -q_{\text{H}_2\text{O}}$$

$$(m) (0.841 \text{ J/g}\cdot^\circ\text{C}) (815.0^\circ\text{C} - 22.3^\circ\text{C}) = - (750.0 \text{ g}) (4.184 \text{ J/g}\cdot^\circ\text{C}) (100.0^\circ\text{C} - 22.3^\circ\text{C})$$

$$m = 405 \text{ g}$$

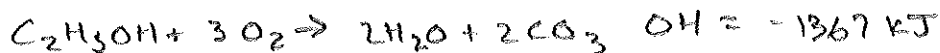
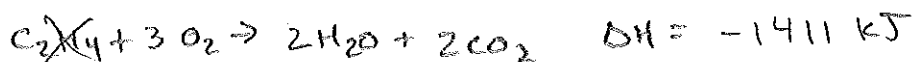
15. Consider the combustion of ethanol, C_2H_5OH :



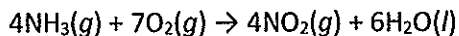
Use Hess' law to calculate ΔH° for this reaction. Equations that may be of use:

- I. $C_2H_4(g) + 3O_2(g) \rightarrow 2H_2O(l) + 2CO_2(g)$ $\Delta H^\circ = -1411 \text{ kJ}$ Keep
 II. $C_{\text{graphite}}(s) + 3H_2(g) + \frac{1}{2}O_2(g) \rightarrow C_2H_5OH(l)$ $\Delta H^\circ = -278 \text{ kJ}$ Do not use
 III. $C_2H_4(g) + H_2O(l) \rightarrow C_2H_5OH(l)$ $\Delta H^\circ = -44 \text{ kJ}$ Flip

- a) 632 kJ
 b) 44 kJ
 c) -1367 kJ
 d) -1742 kJ
 e) -2348 kJ



16. Calculate the standard enthalpy for the following reaction:



Given the following values:

$$\Delta H_f^\circ \text{ for } NH_3(g) = -46 \text{ kJ/mol}$$

$$\Delta H_f^\circ \text{ for } NO_2(g) = 34 \text{ kJ/mol}$$

$$\Delta H_f^\circ \text{ for } H_2O(l) = -286 \text{ kJ/mol}$$

$$\begin{aligned} & \text{Products} - \text{Reactants} \\ & = \left[(4 \text{ mol } NO_2)(34 \text{ kJ/mol}) + (6 \text{ mol } H_2O)(-286 \text{ kJ/mol}) \right] \\ & - \left[(4 \text{ mol } NH_3)(-46 \text{ kJ/mol}) + (7 \text{ mol } O_2)(0 \text{ kJ/mol}) \right] \\ & = -1396 \text{ kJ} \end{aligned}$$

- a) -46 kJ
 b) -206 kJ
 c) -286 kJ
 d) -1396 kJ
 e) Cannot be determined from the information provided

17. It takes 208.4 kJ of energy to remove 1 mol of electrons from an atom on the surface of rubidium (Rb) metal. What is the maximum wavelength of light capable of removing a single electron from an atom on the surface of solid Rb?

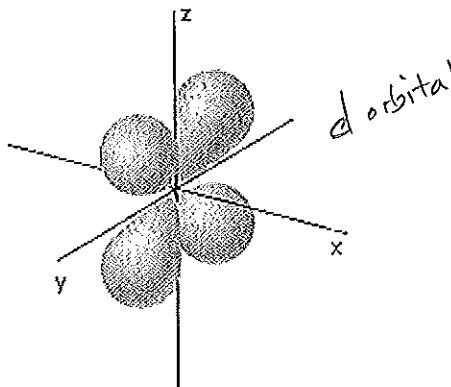
- a) 9.532×10^{-31} nm
- b) 9.532×10^{-11} nm
- c) 9.532×10^1 nm
- d) 5.740×10^2 nm
- e) 9.740×10^5 nm

$$E = 208.4 \text{ kJ} \left(\frac{1000 \text{ J}}{1 \text{ kJ}} \right) \left(\frac{1 \text{ mol}}{6.022 \times 10^{23} \text{ photons}} \right) = 3.46 \times 10^{-19} \frac{\text{J}}{\text{photon}}$$

$$E = \frac{hc}{\lambda} \quad \lambda = \frac{hc}{E} = \frac{6.626 \times 10^{-34} \frac{\text{J}\cdot\text{s}}{\text{photon}} \cdot 3.00 \times 10^8 \frac{\text{m}}{\text{s}}}{3.46 \times 10^{-19} \text{ J/photon}}$$

$$= 5.74 \times 10^{-7} \text{ m} = 574 \text{ nm}$$

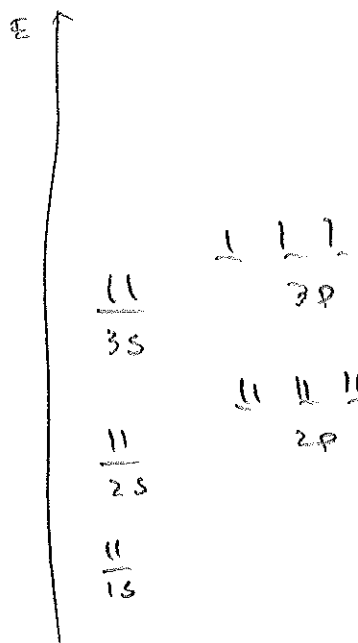
18. The figure below can be labeled:



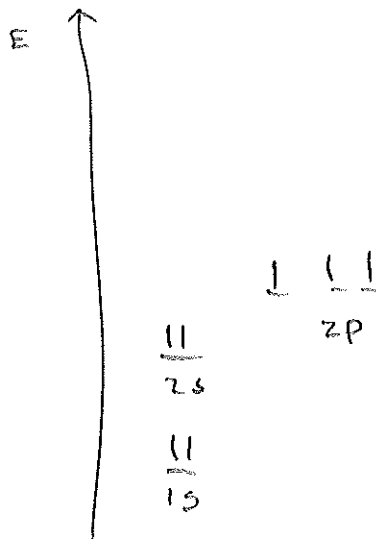
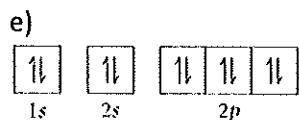
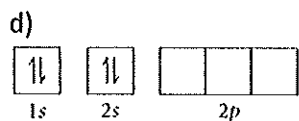
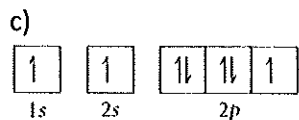
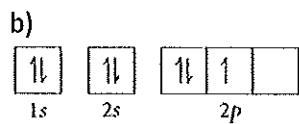
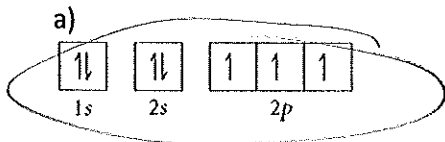
- a) a $2p_x$ atomic orbital
- b) a $2p_z$ atomic orbital
- c) a $5d$ atomic orbital
- d) a $5s$ atomic orbital
- e) a $4s$ atomic orbital

19. Give the complete electronic configuration for P. $\leftarrow 15 e^-$

- a) $1s^2 2s^2 2p^6 3s^1 3d^5$
- b) $1s^2 1p^6 2s^2 2p^5$
- c) $1s^2 2s^2 2p^6 3s^2 3p^3$
- d) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^3$
- e) $1s^2 2s^2 2p^6 3s^2 3d^6$



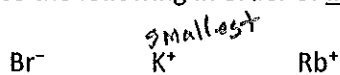
20. Choose the orbital diagram that represents the ground state of N.



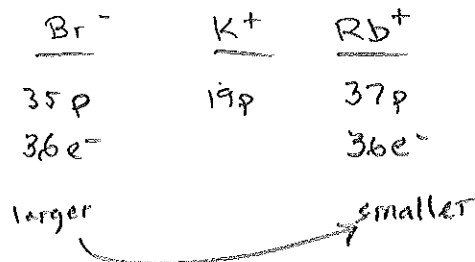
21. Which of the following is **FALSE**?

- a) Na is larger than Na^+
- b) Al is larger than Al^{3+}
- c) O is larger than O^{2-}
- d) O is larger than Ne
- e) Xe is larger than Ne

22. Place the following in order of **increasing** radius.



- a) $\text{Br}^- < \text{Rb}^+ < \text{K}^+$
- b) $\text{K}^+ < \text{Rb}^+ < \text{Br}^-$
- c) $\text{Rb}^+ < \text{Br}^- < \text{K}^+$
- d) $\text{Br}^- < \text{K}^+ < \text{Rb}^+$
- e) $\text{Rb}^+ < \text{K}^+ < \text{Br}^-$



23. Which of the following is FALSE?

- a) Elements in the same group have the same valence electron configuration. T
- b) Atomic radius of elements in a group increases as you go from top to bottom. T
- c) Completely empty and completely full shells are very stable. T
- d) First ionization energy of elements in the same group increases as you go from top to bottom. F
- e) All of the above.

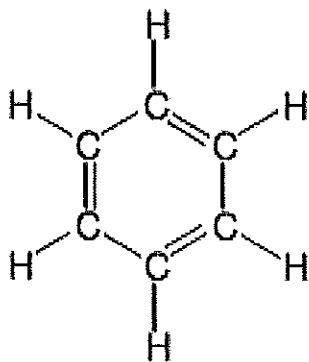
24. Chemical bonds form because:

- a) Doing so lowers the potential energy between the charged particles that compose atoms
- b) Doing so allows the bonded atoms to fit into a smaller volume
- c) Doing so raises the potential energy of the atoms that make up the molecule, which can subsequently release this energy by breaking the bond(s)
- d) Doing so allows the bonded atoms to have a greater density
- e) Doing so allows protons of one atom to combine with electrons of another to form neutrons

25. Under the Lewis model, which best explains why helium (He) obeys the *duet rule* instead of the *octet rule*?

- a) Helium is a noble gas
- b) Helium atoms almost never form covalent bonds
- c) Helium atoms are much less massive than most other elements
- d) The $n = 1$ quantum level fills with only two electrons
- e) None of the above statements are correct

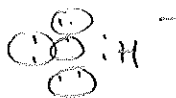
26. All the carbon-carbon bonds in benzene (a Lewis structure is shown below) are known to be the same length. The reason for this is best supported by:



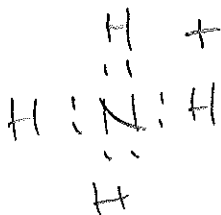
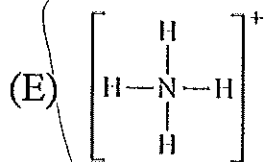
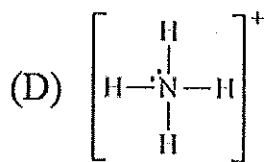
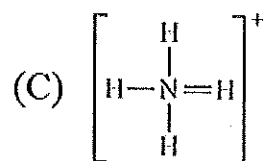
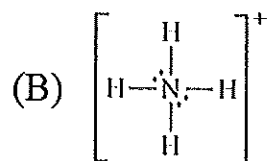
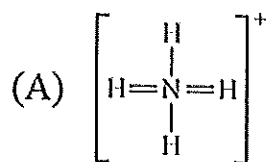
- a) Hund's rule
- b) The concept of resonance
- c) The Aufbau principle
- d) The shape of carbon's 2s orbital
- e) The shape of carbon's 2p orbital

27. The Lewis Dot Structure of the hydroxide ion (OH⁻) depicts:

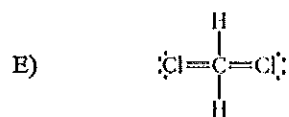
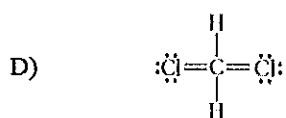
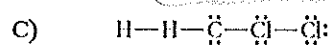
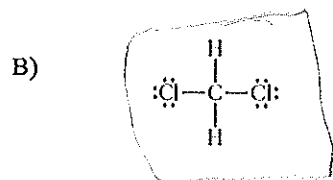
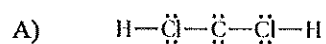
- (A) There are no lone pairs of electrons.
- (B) There is one lone pair of electrons.
- (C) There are two lone pairs of electrons.
- (D) There are three lone pairs of electrons.
- (E) There are four lone pairs of electrons.



28. Choose the best Lewis structure for NH₄⁺.



29. Choose the best Lewis structure for CH_2Cl_2 .



30. The Lewis Dot Structure of H_2O depicts:

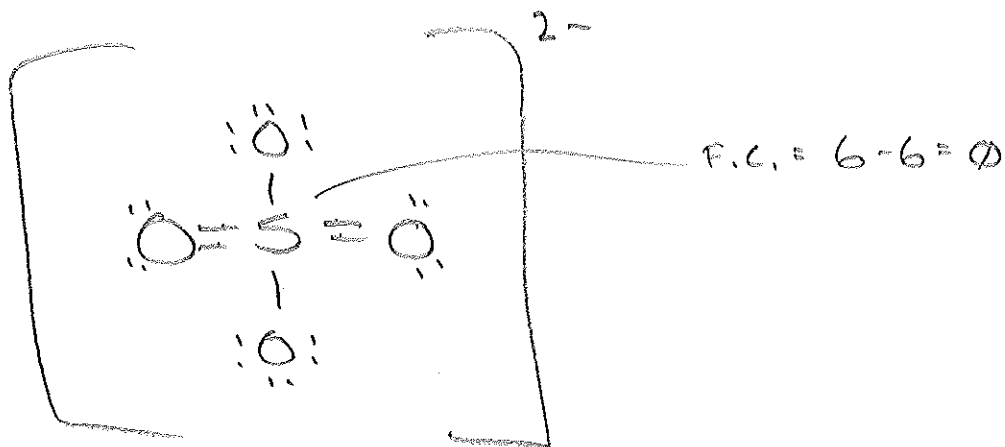
- (A) There are no lone pairs of electrons.
- (B) There is one lone pair of electrons.
- (C) There are two lone pairs of electrons.
- (D) There are three lone pairs of electrons.
- (E) There are four lone pairs of electrons.



[TURN OVER FOR THE FINAL TWO QUESTIONS]

31. A student suggests a Lewis Structure for sulfate (SO_4^{2-}) to have two double bonds and two single bonds to the central sulfur. The formal charge on the sulfur is:

- a) -4
- b) -2
- c) 0
- d) +2
- e) +4



32. Which of the following would not exhibit any resonance structures?

- a) O_3
- b) CF_4
- c) NO_3^-
- d) SO_4^{2-}
- e) CO_3^{2-}

